AULÃO REFFERENCIAL
ENEM - 2016

Prof. Kaiya

I – correta

1 mol $C_8H_{18} \rightarrow 114g$

Calculando o volume de 1 mol do C8H18

1mL $\rightarrow$ 0,75g

x $\rightarrow$ 114g $\Rightarrow$ x = 152mL

Calculando o preço do volume de 1 mol do C_8H_{18}

1L $\rightarrow$ R$ 2,40

0,152L $\rightarrow$ x $\Rightarrow$ x = R$ 0,365

Calculando o preço de refino

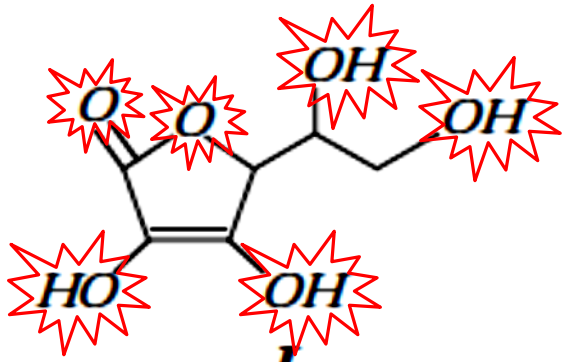
R$ 0,365 $\rightarrow$ 100%

x $\rightarrow$ 31% $\Rightarrow$ x $\cong$ R$ 0,11

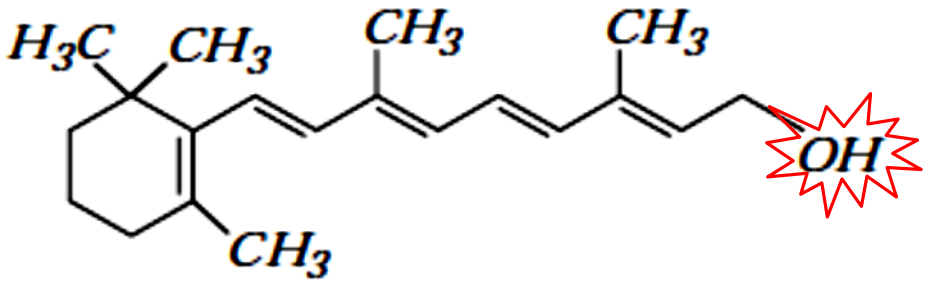
II – incorreta
O nome correto é 2,2,4-trimetilpentano.
III – incorreta

1L de gasolina	→	R$ 2,40	→	100%
		x	→	7%

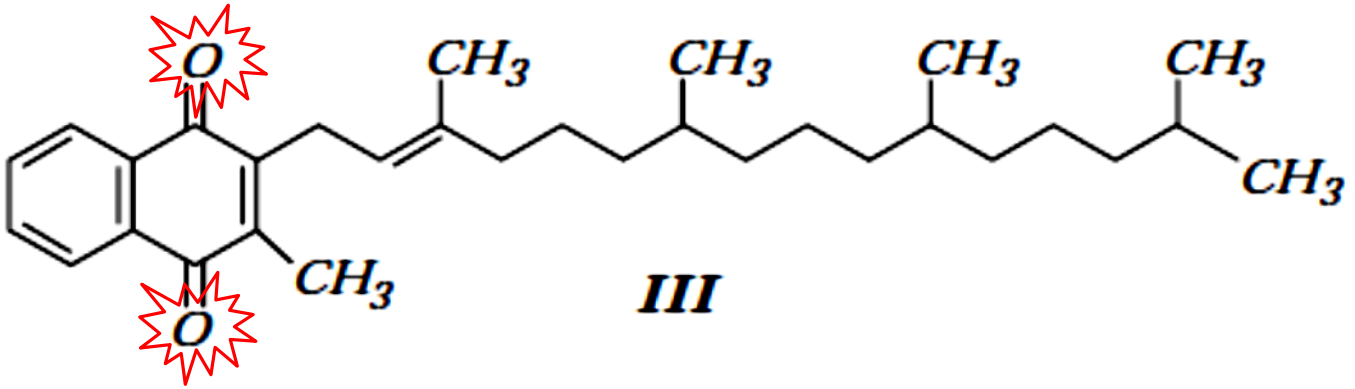
X ≅ R$ 0,17


O
O
OH
OH
HO
OH

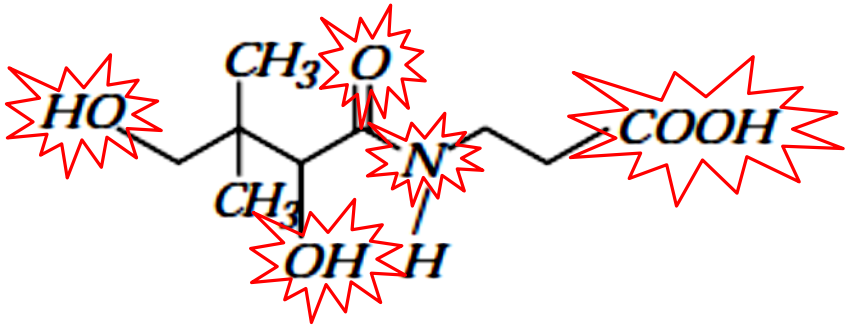
I


H3C
CH3
CH3
CH3
OH
CH3

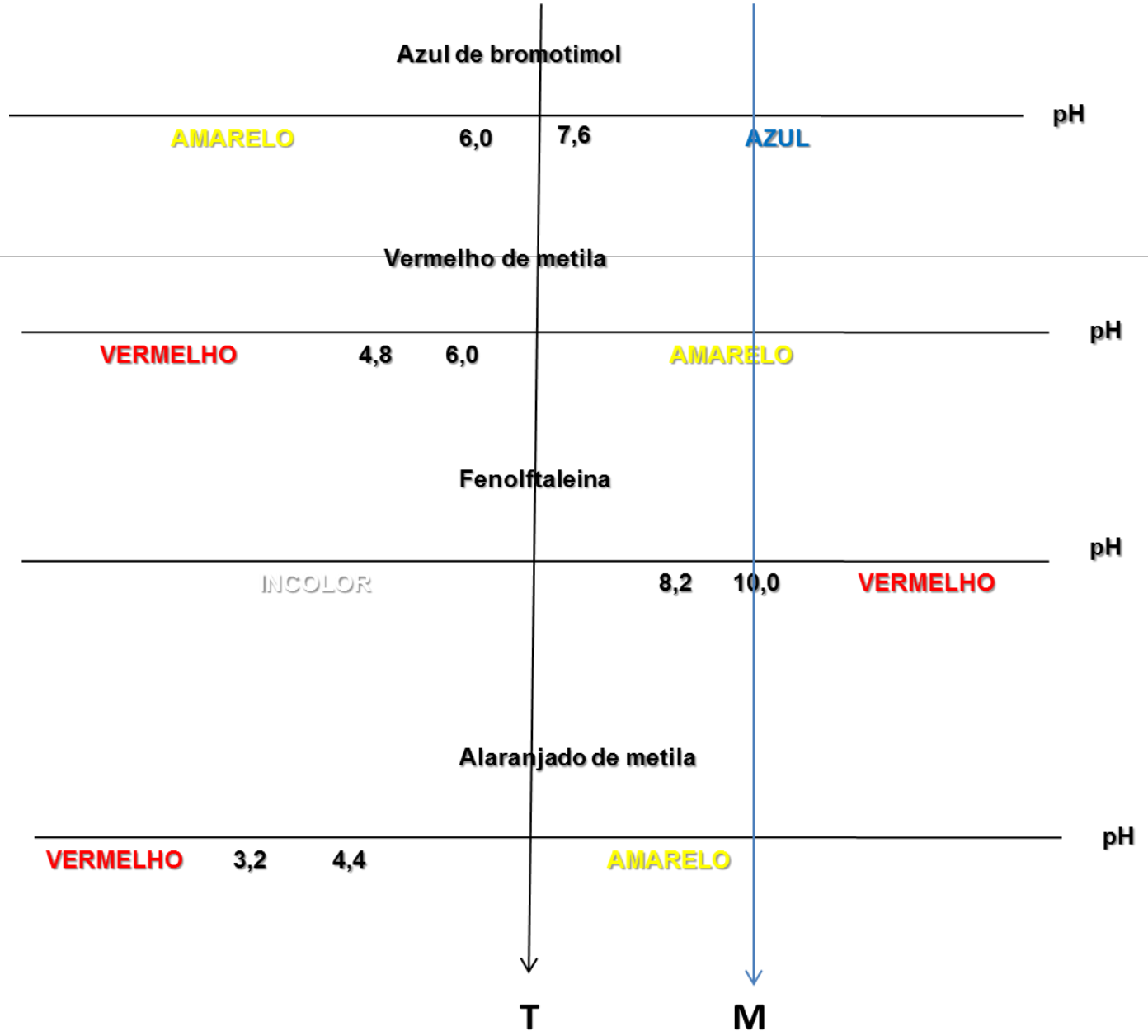
II


O
CH3
CH3
CH3
CH3
CH3
CH3
O

III


HO
CH3
O
N
COOH
CH3
OH
H

IV


Azul de bromotimol
pH
AMARELO
6,0
7,6
AZUL
Vermelho de metila
pH
VERMELHO
4,8
6,0
AMARELO
Fenolftaleina
pH
INCOLOR
8,2
10,0
VERMELHO
Alaranjado de metila
pH
VERMELHO
3,2
4,4
AMARELO
T
M

1 mol de CH_4 $\Rightarrow$ 1 mol de CO_2 $\rightarrow$ - 890 kJ 3

(gás natural)

1 mol de C_4H_{10} $\Rightarrow$ 4 mol de CO_2 $\rightarrow$ -2878 kJ

(GLP) 1 mol de CO_2 $\rightarrow$ x

x = - 719,5 kJ 2

1 mol de C_8H_{18} $\Rightarrow$ 8 mol de CO_2 $\rightarrow$ -5471 kJ

(gasolina) 1 mol de CO_2 $\rightarrow$ x

x = 683,9 kJ 1

BOA PROVA!

Refferencial
COLÉGIO E CURSOS

www.COLEGIOREFFERENCIAL.com.br